

**Figure 1**

**(A)** Schematic diagram illustrating the experimental setup for measuring the effect of temperature on the rate of reaction between hydrogen peroxide and potassium iodide. The reaction mixture is contained in a test tube held by a clamp over a water bath.

**(B)** Graph showing the effect of temperature on the rate of reaction. The y-axis represents Time (s) and the x-axis represents Temperature (°C). The curve shows that as temperature increases, the time taken for the reaction decreases, indicating an increase in the rate of reaction.

**(C)** Graph showing the effect of concentration on the rate of reaction. The y-axis represents Time (s) and the x-axis represents Concentration (mol/L). The curve shows that as concentration increases, the time taken for the reaction decreases, indicating an increase in the rate of reaction.

**(D)** Graph showing the effect of catalyst on the rate of reaction. The y-axis represents Time (s) and the x-axis represents Catalyst Concentration (mol/L). The curve shows that as catalyst concentration increases, the time taken for the reaction decreases, indicating an increase in the rate of reaction.

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